CHAPTER 2

**ATOMIC STRUCTURE AND INTERATOMIC BONDING**

PROBLEM SOLUTIONS

Electrons in Atoms

Problem 2.1

*Cite the difference between atomic mass and atomic weight.*

Answer 2.1

Atomic mass is the mass of an individual atom, whereas atomic weight is the average (weighted) of the atomic masses of an atom's naturally occurring isotopes.

Problem 2.2

*Silicon has three naturally occurring isotopes: 92.23% of 28Si, with an atomic weight of 27.9769 amu, 4.68% of 29Si, with an atomic weight of 28.9765 amu, and 3.09% of 30Si, with an atomic weight of 29.9738 amu. On the basis of these data, confirm that the average atomic weight of Si is 28.0854 amu.*

Solution 2.2

The average atomic weight of silicon  is computed by adding fraction-of-occurrence/atomic weight products for the three isotopes—i.e., using Equation 2.2. (Remember: fraction of occurrence is equal to the percent of occurrence divided by 100.) Thus





Problem 2.3

*Zinc has five naturally occurring isotopes: 48.63% of 64Zn with an atomic weight of 63.929 amu; 27.90% of 66Zn with an atomic weight of 65.926 amu; 4.10% of 67Zn with an atomic weight of 66.927 amu; 18.75% of 68Zn with an atomic weight of 67.925 amu; and 0.62% of 70Zn with an atomic weight of 69.925 amu. Calculate the average atomic weight of Zn.*

Solution 2.3

The average atomic weight of zinc  is computed by adding fraction-of-occurrence—atomic weight products for the five isotopes—i.e., using Equation 2.2. (Remember: fraction of occurrence is equal to the percent of occurrence divided by 100.) Thus



Including data provided in the problem statement we solve for  as



= 65.400 amu

Problem 2.4

*Indium has two naturally occurring isotopes: 113In with an atomic weight of 112.904 amu, and 115In with an atomic weight of 114.904 amu. If the average atomic weight for In is 114.818 amu, calculate the fraction-of-occurrences of these two isotopes.*

Solution 2.4

The average atomic weight of indium  is computed by adding fraction-of-occurrence—atomic weight products for the two isotopes—i.e., using Equation 2.2, or



Because there are just two isotopes, the sum of the fracture-of-occurrences will be 1.000; or



which means that



Substituting into this expression the one noted above for , and incorporating the atomic weight values provided in the problem statement yields









Solving this expression for  yields . Furthermore, because



then



Problem 2.5

*(a) How many grams are there in one amu of a material?*

*(b) Mole, in the context of this book, is taken in units of gram-mole. On this basis, how many atoms are there in a pound-mole of a substance?*

Solution 2.5

(a) In order to determine the number of grams in one amu of material, appropriate manipulation of the amu/atom, g/mol, and atom/mol relationships is all that is necessary, as



= 1.66 × 10−24 g/amu

(b) Since there are 453.6 g/lbm,



= 2.73 × 1026 atoms/lb-mol

Problem 2.6

*(a)* *Cite two important quantum-mechanical concepts associated with the Bohr model of the atom.*

*(b) Cite two important additional refinements that resulted from the wave-mechanical atomic model.*

Answer 2.6

(a) Two important quantum-mechanical concepts associated with the Bohr model of the atom are (1) electrons are particles moving in discrete orbitals, and (2) electron energy is quantized into shells.

(b) Two important refinements resulting from the wave-mechanical atomic model are (1) that electron position is described in terms of a probability distribution, and (2) electron energy is quantized into both shells and subshells--each electron is characterized by four quantum numbers.

Problem 2.7

*Relative to electrons and electron states, what does each of the four quantum numbers specify?*

Answer 2.7

The *n* quantum number designates the electron shell.

The *l* quantum number designates the electron subshell.

The *ml* quantum number designates the number of electron states in each electron subshell.

The *ms* quantum number designates the spin moment on each electron.

Problem 2.8

*For the K shell, the four quantum numbers for each of the two electrons in the 1s state, in the order of nlmlms, are 100 and 100. Write the four quantum numbers for all of the electrons in the L and M shells, and note which correspond to the s, p, and d subshells.*

Answer 2.8

For the *L* state, *n* = 2, and eight electron states are possible. Possible *l* values are 0 and 1, while possible *ml* values are 0 and ±1; and possible *ms* values are  Therefore, for the *s* states, the quantum numbers are  and . For the *p* states, the quantum numbers are , , , , , and .

For the *M* state, *n* = 3, and 18 states are possible. Possible *l* values are 0, 1, and 2; possible *ml* values are 0, ±1, and ±2; and possible *ms* values are  Therefore, for the *s* states, the quantum numbers are , , for the *p* states they are , , , , , and ; for the *d* states they are , , , , , , , , , and .

Problem 2.9

*Give the electron configurations for the following ions: P5+, P3–, Sn4+, Se2–, I–, and Ni2+.*

Solution 2.9

The electron configurations for the ions are determined using Table 2.2 (and Figure 2.8).

P5+: From Table 2.2, the electron configuration for an atom of phosphorus is 1*s*22*s*22*p*63*s*23*p*3. In order to become an ion with a plus five charge, it must lose five electrons—in this case the three 3*p* and the two 3*s*. Thus, the electron configuration for a P5+ ion is 1*s*22*s*22*p*6.

P3*–*: From Table 2.2, the electron configuration for an atom of phosphorus is 1*s*22*s*22*p*63*s*23*p*3. In order to become an ion with a minus three charge, it must acquire three electrons—in this case another three 3*p*. Thus, the electron configuration for a P3*–* ion is 1*s*22*s*22*p*63*s*23*p*6.

Sn4+: From the periodic table, Figure 2.8, the atomic number for tin is 50, which means that it has fifty electrons and an electron configuration of 1*s*22*s*22*p*63*s*23*p*63*d*104*s*24*p*64*d*105*s*25*p*2. In order to become an ion with a plus four charge, it must lose four electrons—in this case the two 4*s* and two 5*p*. Thus, the electron configuration for an Sn4+ ion is 1*s*22*s*22*p*63*s*23*p*63*d*104*s*24*p*64*d*10.

Se2–: From Table 2.2, the electron configuration for an atom of selenium is 1*s*22*s*22*p*63*s*23*p*63*d*104*s*24*p*4. In order to become an ion with a minus two charge, it must acquire two electrons—in this case another two 4*p*. Thus, the electron configuration for an Se2– ion is 1*s*22*s*22*p*63*s*23*p*63*d*104*s*24*p*6.

I–: From the periodic table, Figure 2.8, the atomic number for iodine is 53, which means that it has fifty three electrons and an electron configuration of 1*s*22*s*22*p*63*s*23*p*63*d*104*s*24*p*64*d*105*s*25*p*5. In order to become an ion with a minus one charge, it must acquire one electron—in this case another 5*p*. Thus, the electron configuration for an I– ion is 1*s*22*s*22*p*63*s*23*p*63*d*104*s*24*p*64*d*105*s*25*p*6.

Ni2+: From Table 2.2, the electron configuration for an atom of nickel is 1*s*22*s*22*p*63*s*23*p*63*d*84*s*2. In order to become an ion with a plus two charge, it must lose two electrons—in this case the two 4*s*. Thus, the electron configuration for a Ni2+ ion is 1*s*22*s*22*p*63*s*23*p*63*d*8.

Problem 2.10

*Potassium iodide (KI) exhibits predominantly ionic bonding. The K+ and I– ions have electron structures that are identical to which two inert gases?*

Solution 2.10

The K+ ion is just a potassium atom that has lost one electron; therefore, it has an electron configuration the same as argon (Figure 2.8).

The I– ion is an iodine atom that has acquired one extra electron; therefore, it has an electron configuration the same as xenon.

The Periodic Table

Problem 2.11

*With regard to electron configuration, what do all the elements in Group IIA of the periodic table have in common?*

Answer 2.11

Each of the elements in Group IIA has two *s* electrons.

Problem 2.12

*To what group in the periodic table would an element with atomic number 112 belong?*

Solution 2.12

From the periodic table (Figure 2.8) the element having atomic number 112 would belong to group IIB. According to Figure 2.8, Ds, having an atomic number of 110 lies below Pt in the periodic table and in the right-most column of group VIII. Moving two columns to the right puts element 112 under Hg and in group IIB.

This element has been artificially created and given the name Copernicium with the symbol Cn. It was named after Nicolaus Copernicus, the Polish scientist who proposed that the earth moves around the sun (and not vice versa).

Problem 2.13

*Without consulting Figure 2.8 or Table 2.2, determine whether each of the following electron configurations is an inert gas, a halogen, an alkali metal, an alkaline earth metal, or a transition metal. Justify your choices.*

*(a****)*** *1s22s22p63s23p5*

*(b****)*** *1s22s22p63s23p63d74s2*

*(c)**1s22s22p63s23p63d104s24p6*

*(d****)*** *1s22s22p63s23p64s1*

*(e)**1s22s22p63s23p63d104s24p64d55s2*

*(f)**1s22s22p63s2*

Solution 2.13

(a) The 1*s*22*s*22*p*63*s*23*p*5 electron configuration is that of a halogen because it is one electron deficient from having a filled *p* subshell.

(b) The 1*s*22*s*22*p*63*s*23*p*63*d*74*s*2 electron configuration is that of a transition metal because of an incomplete *d* subshell.

(c) The 1*s*22*s*22*p*63*s*23*p*63*d*104*s*24*p*6 electron configuration is that of an inert gas because of filled 4*s* and 4*p* subshells.

(d) The 1*s*22*s*22*p*63*s*23*p*64*s*1 electron configuration is that of an alkali metal because of a single *s* electron.

(e) The 1*s*22*s*22*p*63*s*23*p*63*d*104*s*24*p*64*d*55*s*2 electron configuration is that of a transition metal because of an incomplete *d* subshell.

(f) The 1*s*22*s*22*p*63*s*2 electron configuration is that of an alkaline earth metal because of two *s* electrons.

Problem 2.14

*(a)**What electron subshell is being filled for the rare earth series of elements on the periodic table?*

*(b)**What electron subshell is being filled for the actinide series?*

Solution 2.14

(a) The 4*f* subshell is being filled for the rare earth series of elements.

(b) The 5*f* subshell is being filled for the actinide series of elements.

Bonding Forces and Energies

Problem 2.15

*Calculate the force of attraction between a Ca2+ and an O2– ion whose centers are separated by a distance of 1.25 nm.*

Solution 2.15

To solve this problem for the force of attraction between these two ions it is necessary to use Equation 2.13, which takes on the form of Equation 2.14 when values of the constants *e* and ** are included—that is



If we take ion 1 to be Ca2+ and ion 2 to be O2–, then *Z*1 = +2 and *Z*2 = −2; also, from the problem statement *r* = 1.25 nm = 1.25 × 10−9 m. Thus, using Equation 2.14, we compute the force of attraction between these two ions as follows:





Problem 2.16

*The atomic radii of Mg2+ and F− ions are 0.072 and 0.133 nm, respectively.*

*(a) Calculate the force of attraction between these two ions at their equilibrium interionic separation (i.e., when the ions just touch one another).*

*(b)**What is the force of repulsion at this same separation distance?*

Solution 2.16

This problem is solved in the same manner as Example Problem 2.2.

(a) The force of attraction *FA* is calculated using Equation 2.14 taking the interionic separation *r* to be *r*0 the equilibrium separation distance. This value of *r*0 is the sum of the atomic radii of the Mg2+ and F− ions (per Equation 2.15)—that is





We may now compute *FA* using Equation 2.14. If was assume that ion 1 is Mg2+ and ion 2 is F− then the respective charges on these ions are *Z*1 = , whereas *Z*2 = . Therefore, we determine *FA* as follows:







(b) At the equilibrium separation distance the sum of attractive and repulsive forces is zero according to Equation 2.4. Therefore

*FR* = − *FA*

= − (1.10 × 10−8 N) = − 1.10 × 10−8 N

Problem 2.17

*The force of attraction between a divalent cation and a divalent anion is 1.67 × 10-8 N. If the ionic radius of the cation is 0.080 nm, what is the anion radius?*

Solution 2.17

To begin, let us rewrite Equation 2.15 to read as follows:



in which and represent, respectively, the radii of the cation and anion. Thus, this problem calls for us to determine the value of . However, before this is possible, it is necessary to compute the value of using Equation 2.14, and replacing the parameter *r* with . Solving this expression for  leads to the following:



Here and represent charges on the cation and anion, respectively. Furthermore, inasmuch as both ion are divalent means that  and . The value of is determined as follows:





Using the version of Equation 2.15 given above, and incorporating this value of and also the value of given in the problem statement (0.080 nm) it is possible to solve for :





Problem 2.18

*The net potential energy between two adjacent ions, EN, may be represented by the sum of Equations 2.9 and 2.11; that is,*

 *(2.17)*

*Calculate the bonding energy E0 in terms of the parameters A, B, and n using the following procedure:*

*1. Differentiate EN with respect to r, and then set the resulting expression equal to zero, because the curve of EN versus r is a minimum at E0.*

*2. Solve for r in terms of A, B, and n, which yields r*0*, the equilibrium interionic spacing.*

*3. Determine the expression for E0 by substitution of r*0 *into Equation 2.17.*

Solution 2.18

Differentiation of Equation 2.17 yields





Now, solving for *r* (= *r*0)





or



Substitution of this expression for *r*0 into Equation 2.17 and solving for *E* (= *E*0) yields







Problem 2.19

*For a Na+–Cl– ion pair, attractive and repulsive energies EA and ER, respectively, depend on the distance between the ions r, according to*





*For these expressions, energies are expressed in electron volts per Na+–Cl– pair, and r is the distance in nanometers. The net energy EN is just the sum of the preceding two expressions.*

*(a) Superimpose on a single plot EN, ER, and EA versus r up to 1.0 nm.*

*(b) On the basis of this plot, determine (i) the equilibrium spacing r0 between the Na+ and Cl– ions, and (ii) the magnitude of the bonding energy E0 between the two ions.*

*(c) Mathematically determine the r0 and E0 values using the solutions to Problem 2.18, and compare these with the graphical results from part (b).*

Solution 2.19

(a) Curves of *EA*, *ER*, and *EN* are shown on the plot below.



(b) From this plot:

*r*0 = 0.24 nm

*E*0 = −5.3 eV

(c) From Equation 2.17 for *EN*

*A* = 1.436

*B* = 7.32 × 10−6

*n* = 8

Thus,





and



= – 5.32 eV

These values are in excellent agreement with the values determined from the plot in part (b).

Problem 2.20

*Consider a hypothetical X+–Y– ion pair for which the equilibrium interionic spacing and bonding energy values are 0.38 nm and –5.37 eV, respectively. If it is known that n in Equation 2.17 has a value of 8, using the results of Problem 2.18, determine explicit expressions for attractive and repulsive energies EA and ER of Equations 2.9 and 2.11.*

Solution 2.20

(a) This problem gives us, for a hypothetical X+-Y- ion pair, values for *r*0 (0.38 nm), *E*0 (– 5.37 eV), and *n* (8), and asks that we determine explicit expressions for attractive and repulsive energies of Equations 2.9 and 2.11. In essence, it is necessary to compute the values of *A* and *B* in these equations. Expressions for *r*0 and *E*0 in terms of *n*, *A*, and *B* were determined in Problem 2.18, which are as follows:





Thus, we have two simultaneous equations with two unknowns (viz. *A* and *B*). Upon substitution of values for *r*0 and *E*0 in terms of *n*, the above two equations become



and





We now want to solve these two equations simultaneously for values of *A* and *B*. From the first of these two equations, solving for *A*/8*B* leads to



Furthermore, from the above equation the *A* is equal to

 (S2.20)

When the above two expressions for *A*/8*B* and *A* are substituted into the above expression for *E*0 (− 5.37 eV), the following results







Or



Solving for *B* from this equation yields



Furthermore, the value of *A* is determined from Equation S2.20 as follows:





Thus, Equations 2.9 and 2.11 become





Of course these expressions are valid for *r* and *E* in units of nanometers and electron volts, respectively.

Problem 2.21

*The net potential energy EN between two adjacent ions is sometimes represented by the expression*

 *(2.18)*

*in which r is the interionic separation and C, D, and ρ are constants whose values depend on the specific material.*

*(a) Derive an expression for the bonding energy E0 in terms of the equilibrium interionic separation r0 and the constants D and ρ using the following procedure:*

*(i) Differentiate EN with respect to r, and set the resulting expression equal to zero.*

*(ii) Solve for C in terms of D, ρ, and r0.*

*(iii) Determine the expression for E0 by substitution for C in Equation 2.18.*

*(b) Derive another expression for E0 in terms of r0, C, and ρ using a procedure analogous to the one outlined in part* (a).

Solution 2.21

(a) Differentiating Equation 2.18 with respect to *r* yields





At *r* = *r*0, *dE/dr* = 0, and

 (2.18a)

Solving for *C* yields



Substitution of this expression for *C* into Equation 2.18 yields an expression for *E*0 as







(b) Now solving for *D* from Equation 2.18a above yields



Substitution of this expression for *D* into Equation 2.18 yields an expression for *E*0 as







Primary Interatomic Bonds

Problem 2.22

*(a) Briefly cite the main differences among ionic, covalent, and metallic bonding.*

*(b) State the Pauli exclusion principle.*

Answer 2.22

(a) The main differences between the various forms of primary bonding are:

*Ionic*--there is electrostatic attraction between oppositely charged ions.

*Covalent*--there is electron sharing between two adjacent atoms such that each atom assumes a stable electron configuration.

*Metallic*--the positively charged ion cores are shielded from one another, and also "glued" together by the sea of valence electrons.

(b) The Pauli exclusion principle states that each electron state can hold no more than two electrons, which must have opposite spins.

Problem 2.23

*Make a plot of bonding energy versus melting temperature for the metals listed in Table 2.3. Using this plot, approximate the bonding energy for molybdenum, which has a melting temperature of 2617°C.*

Solution 2.23

Below is plotted the bonding energy versus melting temperature for these four metals. From this plot, the bonding energy for molybdenum (melting temperature of 2617C) should be approximately 680 kJ/mol. The experimental value is 660 kJ/mol.



Secondary Bonding or van der Waals Bonding

Problem 2.24

*Explain why hydrogen fluoride (HF) has a higher boiling temperature than hydrogen chloride (HCl) (19.4 vs. –85°C), even though HF has a lower molecular weight.*

Solution 2.24

The intermolecular bonding for HF is hydrogen, whereas for HCl, the intermolecular bonding is van der Waals. Since the hydrogen bond is stronger than van der Waals, HF will have a higher melting temperature.

Mixed Bonding

Problem 2.25

*Compute the %IC of the interatomic bond for each of the following compounds: MgO, GaP, CsF, CdS, and FeO.*

Solution 2.25

The percent ionic character is a function of the electron negativities of the ions *X*A and *X*B according to Equation 2.16. The electronegativities of the elements are found in Figure 2.9.

For MgO, *X*Mg = 1.2 and *X*O = 3.5, and therefore,



For GaP, *X*Ga = 1.6 and *X*P = 2.1, and therefore,



For CsF, *X*Cs = 0.7 and *X*F= 4.0, and therefore,



For CdS, *X*Cd = 1.7 and *X*S = 2.5, and therefore,



For FeO, *X*Fe = 1.8 and *X*O = 3.5, and therefore,



Problem 2.26

*(a) Calculate %IC of the interatomic bonds for the intermetallic compound Al6Mn*

*(b) On the basis of this result what type of interatomic bonding would you expect to be found in Al6Mn?*

Solution 2.26

(a) The percent ionic character is a function of the electron negativities of the ions *X*A and *X*B according to Equation 2.16. The electronegativities for Al and Mn (Figure 2.9) are both 1.5 and. Therefore the percent ionic character is determined using Equation 2.16 as follows:



(b) Because the percent ionic character is zero and this intermetallic compound is composed of two metals, the bonding is completely metallic.

Bonding Type-Material Classification Correlations

Problem 2.27

*What type(s) of bonding would be expected for each of the following materials: solid xenon, calcium fluoride (CaF2), bronze, cadmium telluride (CdTe), rubber, and tungsten?*

Solution 2.27

For solid xenon, the bonding is van der Waals since xenon is an inert gas.

For CaF2, the bonding is predominantly ionic (but with some slight covalent character) on the basis of the relative positions of Ca and F in the periodic table.

For bronze, the bonding is metallic since it is a metal alloy (composed of copper and tin).

For CdTe, the bonding is predominantly covalent (with some slight ionic character) on the basis of the relative positions of Cd and Te in the periodic table.

For rubber, the bonding is covalent with some van der Waals. (Rubber is composed primarily of carbon and hydrogen atoms.)

For tungsten, the bonding is metallic since it is a metallic element from the periodic table.

**FUNDAMENTALS OF ENGINEERING QUESTIONS AND PROBLEMS**

Problem 2.1FE

*Which of the following electron configurations is for an inert gas?*

*(A) 1s22s22p63s23p6*

*(B) 1s22s22p63s2*

*(C) 1s22s22p63s23p64s1*

*(D)**1s22s22p63s23p63d24s2*

Solution 2.1FE

The correct answer is A. The 1*s*22*s*22*p*63*s*23*p*6 electron configuration is that of an inert gas because of filled 3*s* and 3*p* subshells.

Problem 2.2FE

*What type(s) of bonding would be expected for bronze (a copper-tin alloy)?*

*(A) Ionic bonding*

*(B) Metallic bonding*

*(C) Covalent bonding with some van der Waals bonding*

*(D) van der Waals bonding*

Solution 2.2FE

The correct answer is B. For bronze, the bonding is metallic because it is a metal alloy.

Problem 2.3FE

*What type(s) of bonding would be expected for rubber?*

*(A) Ionic bonding*

*(B) Metallic bonding*

*(C) Covalent bonding with some van der Waals bonding*

*(D) van der Waals bonding*

Solution 2.3FE

The correct answer is C. For rubber, the bonding is covalent with some van der Waals bonding. (Rubber is composed primarily of carbon and hydrogen atoms.)